

Soil acid Management using Biochar:Review

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Abstract

Soil acidity affects the growth of crops because acidic soil contain toxic levels of aluminum and manganese and characterized by deficiency of essential plant nutrients such as P,N, K, Ca, Mg, and Mo. The management options to correct acid soils are, improve nutrient use efficiency, liming, application of organic materials, biochar, appropriate crop rotations, crop mixtures, use of varieties that tolerant to Al and Mn toxicity. Biochar application improves crop productivity through enhancing water holding capacity, cation exchange capacity (CEC), and adsorption of plant nutrients and creates suitable condition for soil micro-organisms.

Keywords: Biochar, Management, Soil acidity

Introduction

Soil acidification is a natural process, which can either be accelerated by the activity of plants, animals and humans or can be impeded by sound management practices. Industrial and mining activities can lead to soil acidification through acid rains resulting primarily from pyrite oxidation and from acid precipitation caused by the emission of sulfur (SO_x) and nitrogen (NO_x) gases. In areas that remain unaffected by industrial pollution, soil acidification is mainly caused by the release of H⁺ ions during the transformation and cycling of C, N and S in managed ecosystems.

Soil acidification can exert itself in several ways:

- increase of soil acidity or decrease of pH
- decrease of base saturation
- unbalanced availability of elements in the root environment
- decrease of the acid neutralizing capacity (ANC) of the soil (van Breemen 1991)

Soil acidity affects the growth of crops because acidic soil contain toxic levels of aluminum and manganese and characterized by deficiency of essential plant nutrients such as P,N, K, Ca, Mg, and Mo [1]. The explanation for soil acidity might be the sort of parent materials from which the soil are formed, leaching of base forming cations, continuous use of acid forming fertilizers such as Urea and DAP [2]. Soil acidity is one among the most factors that limit and stop profitable and sustained agricultural productivity in many parts of the planet. It is estimated that approximately 50% of the worlds' arable soils are acidic and may be subjected to the effect of aluminum (Al) toxicity of which the tropics and subtropics account for 60% of the acid soils within the world [3]. In Sub-Saharan African countries, soil fertility depletion is that the fundamental biophysical cause for declining per capita food production [4]. This challenge will continue as population pressure increases and degradation of soil resources is aggravated. Reversing this trend lies in the enhancement of sustainable development of the agricultural sector; one of which is the management of good soil quality. In Ethiopia 40.9% of the soil is acidic; from this 28% by moderately acidic soils (pH 4.5 -5.5) and 13% by strongly acidic soils (pH <4.5) [5]. The management options to correct acid soils are, improve

nutrient use efficiency, increase crop production on acidic soils include liming, application of organic materials, biochar, appropriate crop rotations, crop mixtures and use of plant varieties tolerant to Al and Mn toxicity [6].

The potential of biochar as a soil amendment/soil acid management in agricultural fields is a recently recognized and yet it is underutilized technology. Biochar, (commonly known as charcoal or agrichar) is defined as a carbon(C) rich product derived from the pyrolysis of organic material at relatively low temperatures (<700°C) [7]. Biochar application improves crop productivity through enhancing water holding capacity, cation exchange capacity (CEC), adsorption of plant nutrients and creates suitable condition for soil micro-organisms (Lehmann *et al.*, 2011). Biochar can be produced from different plant materials including wood chip and wood pellets, tree bark, crop residues, grasses, organic wastes [8,9] reported an estimate of 30 to 40% by weight of charcoal production with barbeques system in Colombia is a waste. The wastage varies depending on the charring method, the feedstock type, the environmental conditions, and on post charring management of the charcoal. The fraction of waste from charcoal lies between 10 and 15% [9] According to FAO (20 11), charcoal production in Ethiopia is estimated to be 3.92 million tonnes. From the production of this, a minimum of 4% of the charcoal is non- marketable, 156,800 tonnes of biochar can be obtained. Soil management by biochar increases nutrient availability, pH, CEC, crop yields and decreases risk of crop failure as well as opens new possibilities for cropping of high-value crops [10]. Despite the high potentials of biochar in Ethiopia, only very limited information exists [11]. Therefore, this review is carried out to investigate the effect of biochar on acidic soil.

Review Literature

The Chemistry of Soil Acidity

Soil pH range

On the basis of their relative degree of acidity, soils are divided into several acidity or alkalinity classes, as shown in figure 1. Such a classification enables the uses of proper terms for indicating acid-base conditions in soils (**Figure 1**).

General soil pH ranges and reaction classes of soils Source [12].

Acid soils (pH<7) are common in humid regions. In these soils, the concentration of H⁺ ions exceeds that of OH⁻ ions. Most plants grow best in soils with a slightly acidic reaction. In this pH ranges nearly all plant nutrients are available in optimal amounts. Soils with pH<6 will more likely be deficient of some of available nutrients for optimal plant growth. Calcium, magnesium, and potassium are especially deficient in acidic soils. In strongly and very strongly acidic soils, Al, Fe, Mn may exist in toxic quantities because of their increased solubility. In addition, these elements will react with phosphates to form insoluble phosphates, on phosphate retention and fixation (Kim, 2010). (**Figure 2**) Source [13].

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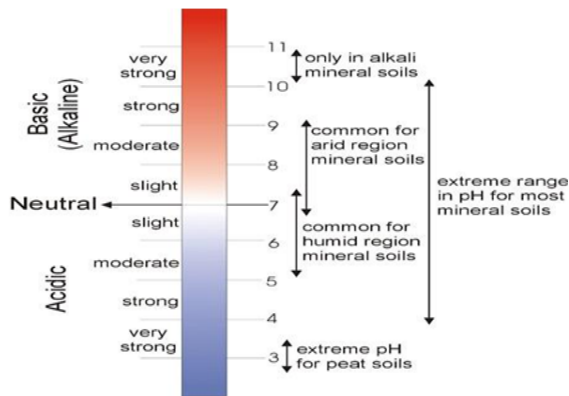


Figure 1. General soil pH ranges and reaction classes of soils.

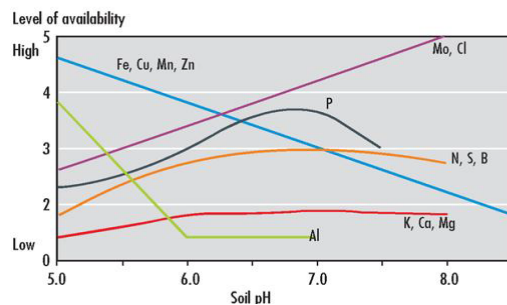


Figure 2. Relation between level of availability of different elements and soil pH.

Active versus potential acidity

A number of compounds contribute to the decomposition of acidic and basic soil reactions. Inorganic/organic acids, produced by the decomposition of organic matter, are common soil constituents that may affect soil acidity. The H^+ ions may be present in soils as adsorbed H^+ ions in the soil solution. The portion of H^+ ions adsorbed by the clay complex becomes exchangeable H^+ ions. The exchangeable H^+ ions dissociate into free H^+ ions, and the amount adsorbed is usually in equilibrium with the amount free in solution. The chemical conditions created produce a chemical environment showing two types of soil acidities [13].

Active acidity is measured and expressed as soil pH. This is the type of soil acidity on which plant growth reacts. However, the tendency exists today to call this exchange acidity salt replaceable acidity which is especially to exchangeable Al^{3+} that can be replaced by an buffered salt solution (e.g. KCl) [13]. However, this is most types of potential acidity because it requires hydrolysis of Al^{3+} to produce the H^+ ions that are the true source for soil acidity. Taken together, the active and potential acidity make up the total soil acidity. The free H^+ ions concentration of the soils solution at any particular time is relatively very small compared to the reserve or exchangeable H^+ concentration.

Residual acidity is defined as soil acidity that remains after active and exchangeable acidity has been neutralized [15]. They believe that residual acidity is associated with aluminum hydroxide ions and with Al^{3+} and H^+ ions that are bound in non-exchangeable forms by organic matter and silicate clays [14] calls this "the remaining soil acidity that cannot be replaced by an unbuffered salts solution".

Exchangeable acidity: Exchangeable hydrogen (H) together with exchangeable aluminum (Al) is known as soil exchangeable acidity. Soil acidity occurs when acidic H^+ ion occurs in the soil solution to a greater extent and when an acid soluble Al^{3+} reacts with water (hydrolysis) and results in the release of H^+ and hydroxyl Al ions into the soil solution [16,17] stated that during soil acidification, protonation increases the mobilization of Al and Al forms serve as a sink for the accumulation of H^+ . The concentration of the H^+ in soils to cause acidity is pronounced at pH values below 4 while excess concentration of Al^{3+} is observed at pH below 5.5 [18] In strongly acidic

conditions of humid regions where rainfall is sufficient to leach exchangeable basic cations, exchangeable Al occupies more than approximately 60% of the effective cation exchange capacity, resulting in a toxic level of aluminum in the soil solution. Generally, the presence of more than 1ppm of Al^{3+} in the soil solution can significantly bring toxicity to plants. Hence, the management of exchangeable Al is a primary concern in acid soils [19].

Causes of soil acidity

- 1) Major reasons for soils to become acidic are: Rainfall and leaching
- 2) Acidic parent material
- 3) Organic matter decay
- 4) Harvest of high yielding crops
- 5) Removal of products from the farm or paddock
- 6) Inappropriate use of nitrogenous fertilizers

The above causes of soil acidity are more easily understood when we consider that a soil is acid when there is an abundance of acidic cations, like hydrogen (H^+) and aluminum (Al^{3+}) present compared to the alkaline cations like calcium (Ca^{2+}), magnesium (Mg^{2+}), potassium (K^+), and sodium (Na^+) [20].

1. Rainfall and leaching: In conditions where rainfall exceeds, leaching during most of the year, the basic soil cations (Ca^{2+} , Mg^{2+} , K^+) are gradually depleted and replaced with cations held in colloidal soil reserves, leading to soil acidity. Soil acidity is really a high rainfall problem [21] Clay soils often contain Fe and hydroxyl Al, which affect the retention and availability of fertilizer cations and anions in acidic soils. Sandy soils are often the first to become acidic because water percolates rapidly, and sandy soils contain only a small reservoir of bases (buffer capacity) due to low clay and organic matter contents. Since the effect of rainfall on acid soil development is very slow, it may take hundreds of years for parent material to become acidic under high rain fall [20].

2. Parent Material: Due to differences in chemical composition of parent materials, soils will become acidic after different lengths of time. Thus, soils that developed from granite material are likely to be more acidic than soils developed from calcareous shale or limestone [20].

3. Organic matter decay /Dissociation: While organic matter has many beneficial effects including improving soil structure, the increasing amount of organic matter may make the soil more acid. Decaying organic matter produces H^+ which is responsible for acidity. The carbon dioxide (CO_2) produced by decaying organic matter reacts with water in the soil to form a weak acid called carbonic acid. This is the same acid that develops when CO_2 in the atmosphere reacts with rain to form acid rain naturally. Several organic acids are also produced by decaying organic matter, but they are also weak acids. Like rainfall, the contribution to acid soil development by decaying organic matter is generally very small, and it would only be the accumulated effects of many years that might ever be measured in a field [22].

4. Crop Production and nutrient removal: Production of all agricultural products causes acidification. The way that plants take up nutrients results in a partitioning of acidity into the soil and alkalinity into the plant as dry matter. If a plant was naturally allowed to die and all parts returned to the soil, there would be no net change in pH. As agriculture removes plant material from a paddock (as grain or pasture) less alkalinity is returned to the soil and the soil becomes more acidic [22].

5. Use of nitrogenous fertilizers: The use of NH_4^+ containing fertilizer over long period leads to soil acidification [23]. However, fertilizers vary in their acidifying effect. The acidifying effect of various N fertilizers followed this order: ammonium sulfate > ammonium nitrate > anhydrous ammonia > urea > calcium nitrate (Table 2). The degree of acidity caused by a fertilizer is modified by soil characteristics, cropping systems and, environmental variables [24]. Fertilizers may also cause acidity by increasing the export of basic cations relative to the unfertilized soil (Table 1) Source: [25]

6. Soil Acidity in Ethiopia: The total area of Ethiopia is 111.8 million hectare out of these only 79 million of hectare is suitable for agriculture. Out of these

Table 1. Acid generating ammonium nitrifying reactions in soil.

Ammonium nitrate	$\text{NH}_4\text{NO}_3 + 2\text{O}_2 \Rightarrow 2\text{NO}_3^- + 2\text{H}^+ + \text{H}_2\text{O}$
Urea	$(\text{NH}_2)_2\text{CO} + 4\text{O}_2 \Rightarrow 2\text{NO}_3^- + 2\text{H}^+ + \text{CO}_2 + \text{H}_2\text{O}$
Anhydrous ammonia.	$\text{NH}_3 + 2\text{O}_2 \Rightarrow \text{NO}_3^- + \text{H}^+ + \text{H}_2\text{O}$
Ammonium phosphate	$\text{NH}_4\text{H}_2\text{PO}_4 + 2\text{O}_2 \Rightarrow \text{NO}_3^- + \text{H}_2\text{PO}_4^- + \text{H}_2\text{O} + 2\text{H}^+$

about 40.9% of area is covered by strongly to weak acid soils. From these 28% moderately to weak acids with pH 5.8-6.7 and 13% covered by strong to moderate acidic soils with pH less than 5.5 [26]. The strongly acid soils are found in ecologies which receive or have historically received high incidence of rainfall and have warm temperatures much of the year. They are often found in Oxisols, Nitisols, and Ferralsols. Thus, the most strongly acidic soils are found in western and south western parts of Ethiopia, the central highlands, the high rainfall areas of north western part of the country. Nevertheless, moderately acidic soils (pH 5.5- 6.5) are distributed through much of the rest of the country [26].

Effect of Soil Acidity

Major effects of soil acidification: Nutrient toxicity can occur in acid soils when the pH is <4.8 or lower. The two most important toxicities in acid soils are those of aluminum (Al) and manganese (Mn) [27]. In strongly acid soils (pH < 4.3) aluminum and manganese become more available in the soil solution and are harmful to plant roots. Aluminum toxicity is the most common and results in decreased production. On plant symptom on acidic soils and causes root stunting. Reduced root growth impedes nutrient and water [28] (**Figure 3**) Source [36].

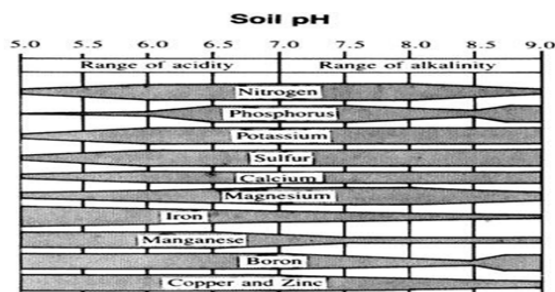
Production and composition of biochar

produced by pyrolysis process

Biochar results from pyrolysis, a thermochemical conversion process for biomass materials such as crop residues, forestry waste, industrial by-product, municipal waste and animal manure. Pyrolysis means breaking chemical bonds (lysis) with heat (pyro). The chemical bonds in biomass break during the pyrolysis process under low oxygen conditions (so the material does not just combust) and different compounds form and recondense into different end products. Together with the cracking reactions responsible for the formation of volatiles (gases and liquids), polycondensation reactions also occur which form biochar. Pyrolysis generally generates three products: a gas, a liquid, and a solid. Biochar is the solid [29].

Chemical composition of biochar (Table 2)

Source: [30]

**Figure 3.** Effect of pH on nutrient availability.**Table 2.** Inorganic properties of biomass samples (wt.% of the ash).

Sample	SiO ₂	C	Al ₂ O ₃	N	Fe ₂ O ₃	Ash	CaO	MgO	Na ₂ O	K ₂ O	SO ₃	P ₂ O ₅
Olive husk	32.9	50.2	8.4	1.0	6.3	4.1	14.5	4.2	26.4	4.3	0.6	2.5
Corn cob	52.9	49.0	9.1	0.5	6.8	1.1	9.4	3.3	1.6	5.3	4.9	6.6
Tea waste	44.4	48.2	7.2	0.5	5.2	1.5	12.6	3.8	1.8	18.8	1.4	4.8

Uses of Biochar

The main purpose for the preparation of biochar is for carbon sequestration (stores carbon for long time), improved soil fertility, and reduces soil acidity for better crop production (IBI, 2012). It improves crop yield when applied as a soil amendment. Biochar can also be burned for energy and processed into activated charcoal for water purification [10].

Methodology

Description of the Study Area

The study was carried out at Koga which is Mecha District, West Gojam Zone of Amhara National Regional state, Ethiopia. Nitisol is the dominant and soil type is moderately acidic [14].

Experimental Setup

Three rates of biochar (4, 8, 12 t ha⁻¹), one rate of lime (2 t ha⁻¹) and a control treatment (without biochar and lime) were combined in a factorial arrangement with three rates of chemical N and P fertilizer rates (0, ½ and full recommended). A randomized complete block design (RCBD) with three replications was used [14].

Eucalyptus char from commercial sellers at Merawi town was collected and the bigger sized were crushed and fine grained in order to distribute them uniformly and increase the surface area of the char for better soil-char contact (**Figure 2**). Lime produced from Dejen lime mill enterprise was used [14] (**Figure 4**).

Soil Sampling

Prior to planting and before application of biochar and lime, composite soil samples were taken from the experimental sites using soil auger in 0-20 cm soil depth. At harvesting, soil samples from each treatment were taken and independently analyzed. Soil samples were air dried under shade, ground by mortar & pestle and sieved to pass through 2 mm mesh [14].

Soil Analysis

The soil pH was determined using glass electrode pH meter in 1:2.5 soil to water ratio in volume and exchangeable acidity (Al³⁺ and H⁺) were extracted with KCl solutions [29]. Total nitrogen analysis was done using the Kjeldahl method [31]. Available phosphorus was determined by Bray II method [32]. Cation exchange capacity (CEC) was determined using the ammonium replacement method. Soil Organic Carbon (SOC) was determined following the wet digest ion method used by Walkley and Black [33] and organic carbon was converted to organic matter by multiplying with a coefficient of 1.724.

Result and Discussion

Effect of Biochar on Soil Properties

Soil pH and Exchangeable Acidity: Addition of biochar from lower rate to higher rates production a reduction tends of exchangeable acidity. Incorporation of 12 t ha⁻¹ biochar gave the lowest exchangeable acidity (0.39 cmol kg⁻¹) followed by 2 t ha⁻¹ lime (0.44 cmol kg⁻¹), where the highest exchangeable acidity was (0.60 cmol kg⁻¹) from the control plots (**Figure 3**) [14]. Lowering of exchangeable acidity and a rise in pH can provide a wide range of benefits in terms of soil quality, not only by chemically improving the availability of plant nutrients, and in some cases by reducing the availability of detrimental elements such as Al [34] (**Figure 5**).



Figure 4. The state of charcoal before (left) and after (right) crushed.

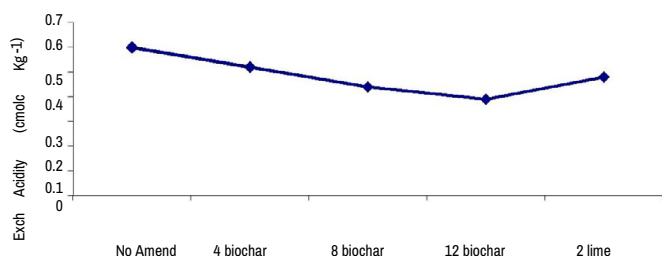


Figure 5. Effect of biochar on exchangeable acidity.

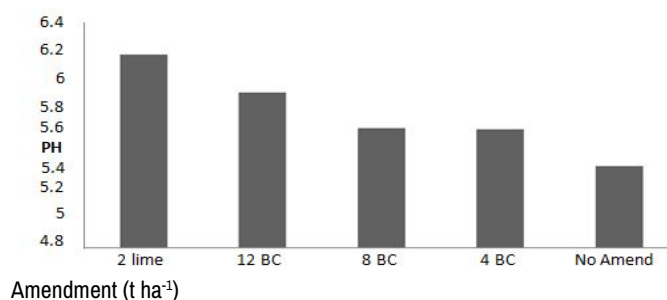


Figure 6. Effect of biochar and lime on soil pH.

Despite the long term potential impacts are to be seen, application of biochar and lime raised the soil pH substantially in one season. The higher pH values (6.17 and 5.9) were obtained on application of 2 t ha⁻¹ lime and 12 t ha⁻¹ biochar, respectively, whereas the lowest pH (5.38) was found on the control plot (Figure 4). Within one season pH increased by 0.79 and 0.52 units by application of 2 t ha⁻¹ lime and 12 t ha⁻¹ biochar, respectively from the control. Application of biochar at rates of 8 and 4 t ha⁻¹ also showed 0.31 and 0.29 unit increment, respectively on soil pH from the same. It was apparent that the biochar applied in this study had a pH value of 8.43, which is basic in reaction [14]. The alkalinity of most biochar can be beneficial to acidic soils, acting as a liming agent to increase pH, and decrease exchangeable Al [10]. (Figure 6).

Available P

Soil available P increased after application of biochar and lime. As biochar rates increased, availability of P also increased. It increased from 12.75 ppm without amendment to 18.92 and 17.50 ppm after application of 12 t ha⁻¹ biochar and 2 t ha⁻¹ lime, respectively (Figure 7).

An increase of available P on application of biochar could happen due to two reasons: the first one is that liming effect of biochar on acid soil that precipitates Al and Fe as Fe(OH)₃ and Al(OH)₃, thus increases availability of phosphorus in the soil system [35]. According to [36] and under acid conditions, phosphorus is precipitated as Fe or Al phosphates. Maximum availability of phosphorus generally occurs in a pH range of 6.0 to 7.0. [37] also indicated that phosphorus is most available near pH 6.5 for mineral soils. [35] also discussed that P availability is at maximum in the pH range 5.5-6.0. Maintaining a soil pH in this range also favors the presence of H₂PO₄⁻ which is more readily absorbed by the plant than HPO₄²⁻ and occur at pH values above 7.0 [35]. Source: (Abewael *et al.*, 2013).

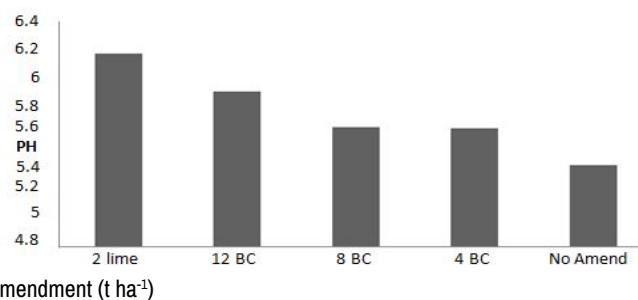


Figure 6. Effect of biochar and lime on soil pH.

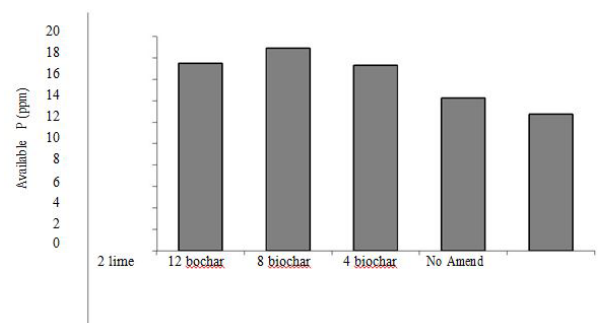


Figure 7. Effect of biochar and lime on available phosphorus.

The second reason is the presence of available phosphorus in biochar itself. (Deluca *et al.*, 2009) reported that biochar inherently contains a high content of soluble P salts formed during the charring of organic materials [38] reviewed that an increase in availability of phosphorus after application of biochar could be due to the direct nutrient addition by the biochar and changes in soil microbial dynamics. (Chan *et al.*, 2007) and (Lehman *et al.* 2006) also discussed that application of biochar increased phosphorus concentration. (Figure 8) Source: [11]

The second reason is the presence of available phosphorus in biochar itself. (Deluca *et al.*, 2009) reported that biochar inherently contains a high content of soluble P salts formed during the charring of organic materials [38] reviewed that an increase in availability of phosphorus after application of biochar could be due to the direct nutrient addition by the biochar and changes in soil microbial dynamics.

Soil analysis result showed that application of fertilizer increased available P; 14.70 ppm, 16.35 and 17.30 ppm were found on the application of 0, 23/30 and 46/60 (N/P₂O₅) kg ha⁻¹ Table 3. Accordingly, application of biochar and lime with fertilizer and without fertilizer showed different values on available P. Biochar rates and lime showed higher P at higher rates of fertilizer. As the rate of biochar and the rate of fertilizer increases, the P value also increases. (Table 3)

This result indicates that the availability of the applied P fertilizer could be supported by the application of biochar and lime.

Cation Exchange Capacity (CEC)

The soil analysis result showed that incorporation of biochar and lime increased CEC of the soil. High CEC values (20.73 and 19.05 cmol kg⁻¹) were found on plots treated with 2 t ha⁻¹ of lime and 12 t ha⁻¹ of biochar, respectively, whereas the lowest CEC (14.69 cmol kg⁻¹) was found in the control plots (Figure 6). Generally, all biochar treatments showed higher CEC than the control one. Application of 4, 8 and 12 t ha⁻¹ of biochar gave 6.13%, 11.25% and 29.68% CEC increment, respectively over the control. Edmunds (2012) also mentioned that biochar has high CEC [11].

Organic Carbon and Total Nitrogen

Based on the soil analysis result, higher organic carbon (1.91% and 1.84%) were found on 2 t ha⁻¹ lime and 12 t ha⁻¹ biochar, respectively whereas the lowest content (1.66%) was found from control plots (Table 1). The soil analysis data

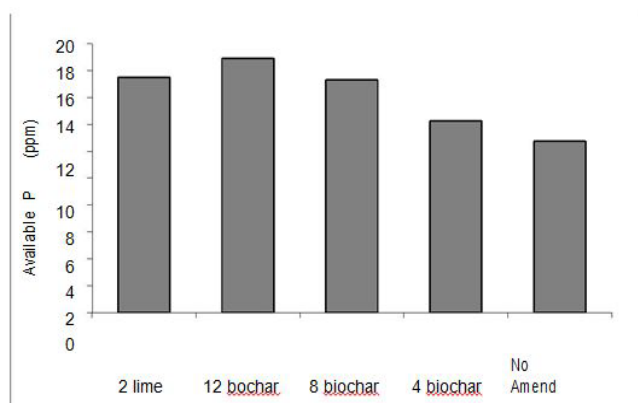


Figure 8. Effect of biochar and lime on available phosphorus.

Table 3. Effect of fertilizer on selected soil chemical properties.

Fertilizer N/P ₂ O ₅ (kg ha ⁻¹)	PH	Total N (%)	Available P (ppm)
0	5.73	0.17	14.70
23/30	5.77	0.18	16.35
46/60	5.75	0.18	17.40

Table 4. Effect of biochar and lime on soil chemical properties.

Amendment	OC (%)	Total N (%)	
2 t ha ⁻¹	Lime	1.91	0.19
12 t ha ⁻¹ biochar	1.84	0.18	
8 t ha ⁻¹	Biochar	1.77	0.18
4 t ha ⁻¹	Biochar	1.70	0.17
No Amendment	1.66	0.17	

revealed that soil organic carbon increased by biochar application. This is due to the raise of soil pH in short period of time that favors soil microbes to decompose crop residues (Table 4).

Application of biochar resulted in little differences on total nitrogen. Actually the study area has low total nitrogen [11] Intensive cultivation as well as less crop residue management may contribute significant roles for the low level of total nitrogen. Application of biochar has no direct effect on nitrogen content of the soil, but it retains nitrogen from leaching [31]. The review of (Clough *et al.*, 2013) indicated that biochar has a high adsorption capacity of NO₃⁻ that prevents nitrogen leaching from the top soil. [39] concluded that NO₃⁻ was only weakly adsorbed to biochar, that it could be desorbed by water infiltration, and that the net result may be an increase residence time for NO₃⁻ in the soil. Due to its resistance to decomposition in soil, single applications of biochar can provide beneficial effects over several growing seasons in the field (Major *et al.*, 2010). Therefore, biochar does not need to be applied frequently, as is usually the case for manures, compost, and synthetic fertilizers [39].

Yield Response to Application of Biochar, Lime and Mineral Fertilizers

Significant teff yield variation was obtained by incorporating biochar and lime. The highest rate of biochar (12 t ha⁻¹) resulted in better yield, while application of 2 t ha⁻¹ lime resulted in a yield above application of 8 t ha⁻¹ biochar and less than 12 t ha⁻¹ biochar applications (Table 3). In general, biochar and lime applications resulted in a significant yield increase (P<0.01) (Table 3). Application of 12 t ha⁻¹ biochar and 2 t ha⁻¹ lime gave 2.67 and 2.45 t ha⁻¹ grain yield, respectively, while the lowest yield (1.44 t ha⁻¹) was obtained from control plots. Incorporation of 12 t ha⁻¹ biochar, 2 t ha⁻¹ lime, 8 t ha⁻¹ and 4 t ha⁻¹ biochars had 85.66%, 70.63%, 37.79% and 19.97% yield advantage, respectively over the untreated plots.

Similar to grain yield, significant dry biomass yield variation was obtained by incorporation of biochar and lime. The highest rate of biochar (12 t ha⁻¹) had a better biomass yield response than the lower rates. Application of 2 t ha⁻¹ lime resulted in better yield than 8 t ha⁻¹ but less than 12 t ha⁻¹ biochar. Application of 12 t ha⁻¹ biochar and 2 t ha⁻¹ lime gave 17.77 and 15.03 t ha⁻¹ teff dry biomass yield, respectively whereas the lowest biomass yield (11.55 t ha⁻¹) was obtained from control treatment [40].

Conclusions

Application of biochar resulted in positive effect on soil chemical properties/soil acidic management. In highly acidic soils it may serve as a soil amendment by increasing the soil pH, P availability, CEC and SOC in a sustainable manner. Biochar can enhance teff yield through improving the chemical and biological properties of the soil including organic carbon and other plant nutrients. This result also indicates application of biochar alone and in combination with fertilizer significantly increases yield. Significant yield increase was also achieved by combined application of biochar with chemical fertilizer, implying the role of biochar to improve fertilizer use efficiency.

References

1. Tisdale S, Nelson W and Beaton J. "Soil Fertility and Fertilizers". Macm Pub Comp New York (1985) 13: 9780024208309.
2. Cooke G. "Fertilizer for maximum yield". Collins Professional and Technical Books, London, UK (1982).
3. Sumner M and Noble A. "Soil acidification: The world story". Handbook of soil acidity, Marcel Dekker, Inc. New York, NY, USA (2003) 128.
4. Sanchez P.A, Sheperd S., Buresh A.M.N., Izac A.U., Mokwunye, et al." Soil Fertility Replenishment in Africa. An Investment in Natural Resources Capital. "(1997) 1-46.
5. Mesfin A. "Nature and Management of Acid Soils in Ethiopia". Addis Ababa, Ethiopia (2007).
6. Sanchez P and Salinas J.G. "Low input Technology for managing Oxisols and Ultisols in Tropical America. Adv. Agron (1981) 34:279-406.
7. In Lehmann J and Joseph S. " Biochar effects on soil nutrient transformations". Biochar for Envi Manage Sci. tech 251-270.
8. Yaman S. "Pyrolysis of biomass to produce fuels and chemical feed stocks. Energy conversion and management". (2004) 45: 651-671.
9. Lehmann J and Rondon M. "Bio-char soil management on highly-weathered soils in the humid tropics." Biological approaches to sustainable soil systems (2006) 517-530.
10. Major J, Rondon M, Molina D, Riha Sand Lehmann J, et al. Maize yield and nutrition during 4 years after biochar application to a Colombian Savanna oxisol. Plant and soil (2010) 333: 117-128.
11. Abewa A, Yitaferu B., G. Selassie Y. and Amare T. Role of Biochar on Acid soil Reclamation and Yield of Teff (Eragrostis tef [Zucc] Trotter) in Northwestern Ethiopia (2013).
12. Brady N. C and Weil R. R. " The Nature and Properties of Soils". New Jersey: Pearson Prentice Hall (2008).
13. Goedert W, Lobato and S. Lourenco. " Nutrient use efficiency in Brazilian acid soils: Nutrient management and plant efficiency." In: Plant-Soil Interactions at Low P. H. Moniz, A.C. et al. Braz Soil Sci Socie (1997) 97-104.
14. Kim H T. "Principles of soil environmental chemistry. " CRC press New York (2010).
15. Brady N.C and Weil R. "The nature and properties of soils." Prentice Hall Inc (1996) 449 -453.

16. Brady Nand Weil R. "The nature and properties of soils." Prentice-Hall Inc New Jersey, USA (2002) 960.
17. Foth H and BELLIS. "Soil fertility." Lewis CRC Press LLC, USA (1997) 290.
18. Nair K and G Chamuah. "Exchangeable aluminum in soils of Meghlaya and management of Al³⁺ related productive constraints." J. Indian Soc. Soil Sci (1993) 4: 331-334.
19. Buol S, Hole and McCracken. "Soil genesis and classification." IA: The Iowa State University Press. Xiv, New Delhi (1989) 446.
20. Jonhson C. "Cause and effects of soil acidity". Oklahoma Cooperative Extension Facts sheets (1914) 2239.
21. Slattery B and Hollier C. "The Impact of Acid Soils in Victoria". Report for the Department of Natural Resources and Environment, Goulburn Broken Catchment Management Authority, North East Catchment Management Authority (2002).
22. MUnkovich P., Sanford and Pate M. Effects of grazing on plants and soil nitrogen relations of pasture-crop rotations. J Agric (1998) 49:475-485.
23. Bolan N, Hedley T and White. "Process of soil acidification during nitrogen cycling with emphasis on legume based pastures". Plant and soil (1991) 134:53-63.
24. Adams F. "Soil acidity and liming. American Society of Agronomy", Madison, WI (1984).
25. Kennedy IR. "Acid Soil and Acid Rain". (1992).
26. Sahilemedhin S and Taye B. "Procedures for Soil and Plant Analysis Technical Paper." National Soil Research Center (2000)
27. Slattery W, Ridley AM and Windsor SM. "Ash alkalinity of animal and plant products." AUST J EXP AGR (1999) 31: 321-324.
28. Slattery W and Coventry D. "Response of wheat, triticale, barley, and canola to lime on four soil types in north-eastern Victoria." Aust. J. Exp. Agric (2000) 33:609-618.
29. Hunt J, Dupont M, Sato D and Andrew K. "The Basics of Biochar; A Natural Soil Amendment". Soil and Crop Manage-SCM (2010) 30:1-6.
30. Ayhan D. "Effects of temperature and particle size on bio-char yield from pyrolysis of agricultural residues." Department of Chemical Engineering, Selcuk University, 42031 Konya, Turkey (2004).
31. Jackson M. "Soil chemical analysis". New Jersey: Prentice Hall, Inc (1958) 183-204
32. Bray R and Kurtz. "Determination of total, organic and available form of phosphorous in soils. Soil Science". (1945) 59: 39-45.
33. Chan K, Van Zwieten I, Meszaros I, Downie A and Joseph S, et. al. "Agronomic values of green waste biochar as a soil amendment." Aust. J. Soil Res (2007) 45: 629-634.
34. Brady N.C. "The nature and properties of soil." New York (1980).
35. Tisdale S, Nelson W, Beaton J and Havlin J. "Soil fertility and fertilizers". Prentice-hall, inc., Englewood Cliffs, N.J., USA. van Breemen, N. (1991). Soil acidification and alkalinisation. In "Soil Acidity" (2002) 1-7.
36. Griffith B. "Essential role of phosphorus in plants. Retrieved December". (2006).
37. Miller R and Donhue R. "Soil in our environment". Prentice-hall (1997).
38. Lehmann J, Gaunt J and Rondon M. "Bio-char sequestration in terrestrial ecosystems-a review. Mitigation and adaptation strategies for global change". (2006) 11: 395-419.
39. Kameyama K, Miyamoto T, Shiono T and Shinogi Y. "Influence of sugarcane bagasse derived biochar application on nitrate leaching in calcareous dark red soil." Environ. Qual (2012) 41: 1131-1137.
40. Lehmann J, Rillig M. C, Thies J, Masiello C, Hockaday W, et al. "Biochar effects on soil biota: A review. Soil Biol." Soil Chem (2011) 43: 1812-1836.

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